



Research Article

QSAR STUDY HIV-1 PROTEASE (RP) INHIBITORS OF CYCLOALKYLPYRANONE DERIVATIVES

S.S. Tomar¹, Nadreen Banu¹, Kaleem Ahmad^{2*}

¹Department of chemistry R.S.S (P.G.) College Pilkhuwa, Hapur (U.P.) , India

²Department of chemistry Mahila (P.G.) College, Bahraich (U.P.) , India

*Author for correspondence: E-mail: drkaleem2012@gmail.com

Tel. / Fax +91-5252-234179, Mobile No. +91-8004180005

Received: 10 March 2015 / Revised: 23 April 2015 / Accepted: 27 April 2015/ Online publication: 1 July 2015

ABSTRACT

The implications for the ability of quantum chemical descriptors of highest occupied molecular orbital energy, lowest unoccupied molecular orbital energy, electron affinity ionization potential absolute hardness, Softness , Chemical Potential Molar refractivity (MR), Molar Volume (MV) and Parachor (Pc) to describe the biological activities HIV-1 Protease (RP) Inhibitors of Cycloalkylpyranone derivatives are discussed. The relationships between observed biological activity, log₁₀C, and the quantum chemical descriptors are established. For this establishment various QSAR models have been developed. The parameter adopted in this calculation is the semi-empirical PM3 based. The QSAR model sixth provides a good arrangement between Obs log₁₀c & predicted activity.

Key words: Absolute hardness; Chemical potential; Electronegativity; Global Softness; refractivity (MR) , Molar Volume (MV), HOMO; LUMO, Parachor (Pc). PM3

1. INTRODUCTION

The World Health Organization has warned about the danger of AIDS, which has killed more than 2.5 million people worldwide.^{1,2} Hagen and coworkers¹ have reported on advancements in the treatment of HIV infection by the use of HIV protease inhibitors. Originally referred to as HTLV-III or LAV, this enveloped single-stranded RNA virus is now called human immunodeficiency virus (HIV)^{3, 4} and two genetically distinct subtypes, HIV-1 and HIV-2, have been characterized,⁵⁻⁶ of which the former has been found to be prevalent in causing the disease.

In the present study we have taken structures of a set of HIV-1 Protease (RP) Inhibitors of Cycloalkylpyranone derivatives of anti-HIV drugs derivatives and then compared to the numerical values of a biological activity. The challenge here has been to find some numerical information for a molecule. This structure information and the measured property or activities are then converted into a mathematical model of relationship. From a quality model it is possible to predict and to design compounds for synthesis and testing that have a good possibility for activity. In this paper, the multi linear regression analysis has been applied for QSAR study. The relationship has been

worked out between the Log 1/C values of a series of compounds and certain quantum chemical descriptors.

2. MATERIAL AND METHODS

The compounds taken for study are derivatives of a set of HIV-1 Protease (RP) Inhibitors of Cycloalkylpyranone derivatives of anti-HIV drugs shown in Fig.1.

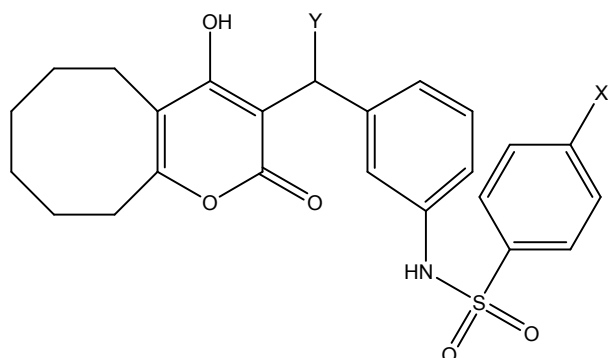


Fig.1: PR Inhibition derivatives of Cycloalkylpyranone

The Quantum Chemical parameter based QSAR study was performed by the following important descriptors like Eigen value of Highest occupied molecular orbital (EHOMO), Eigen value of lowest unoccupied molecular orbital (ELUMO)⁷, Absolute Hardness (η)⁸, Chemical Potential (μ)⁹, Global Softness (S)¹⁰, Electronegativity (χ)¹¹, Molar refractivity (MR)¹², Molar Volume (MV)¹³, Parachor (Pc)¹⁴. The molecules were drawn by spartan06v110, software and the geometries were optimized at PM3 level in conjunction with molecular mechanics. The global hardness and electronegativities were calculated using frontier orbital energies obtained from PM3 results and reported in table 2. Multiple linear regression analysis (MLR) is performed to establish the QSAR. Of HIV-1 Protease (RP) Inhibitors of Cycloalkylpyranone derivatives compounds were taken with their observed activity is shown in table 1.

Table 1. Structural detail and biological activity for the compounds used in the present study

Compound No.	X	Y	Observed log 1/C
1	4-Cl	Cy-C ₃ H ₅	8.600
2	4-Cl	C ₂ H ₅	8.460
3	4-Cl	C ₃ H ₇	8.400
4	4-Cl	CHMe ₂	7.750
5	4-CN	Cy-C ₂ H ₅	9.100
6	4-CN	C ₂ H ₅	8.510
7	4-CN	C ₃ H ₇	8.850
8	4-CN	C ₄ H ₉	8.680
9	4-CN	CHMe ₂	8.230
10	4-CN	CH ₂ CHMe ₂	8.820
11	4-F	Cy-C ₃ H ₅	8.510
12	4-F	C ₃ H ₇	8.680
13	4-F	C ₄ H ₉	8.220
14	4-F	CHMe ₂	7.960
15	4-F	CH ₂ CHMe ₂	8.550

3. RESULTS AND DISCUSSION

3.1 Multiple Linear Regression (MLR) analysis

MLR analyses were performed using Minitab 16 software. The quantum mechanical descriptors were used as independent variables and the Obsd log₁/C₅₀ values as the dependent variables. In the statistical analyses, the systematic search was performed to determine the significant descriptors. The correlation matrix was developed to minimize the effect of co-linearity and to avoid dependencies between subsets of the variables and multi-co-linearity (high multiple correlations between subsets of the variables). The MLR equations of different QSAR models are as follows-

3.2 First QSAR model

MLR equation of this QSAR model P log 1/C is given by-

$$\text{Obsd log } 1/K_i = 5.89 - 1.96 \text{ E LUMO (e.v)} \quad S = 0.523851$$

$$S = 0.162964$$

$$\text{PRESS} = 0.465827$$

$$r^2 = 79.4\%$$

3.3 Second QSAR model

MLR equation of this QSAR model P log 1/C is given by-

$$\text{Obsd log } 1/C = 7.16 - 4.87 \text{ E LUMO (e.v)} + 0.706 \text{ E HOMO (e.v)}$$

$$S = 0.506241$$

$$\text{PRESS} = 10.2998$$

$$r^2 = 87.5\%$$

3.4 Third QSAR model

MLR equation of this QSAR model P log 1/C is given by-

$$\text{Obsd log } 1/C = 34.1 - 6.51 \text{ E LUMO (e.v)} + 2.36 \text{ E HOMO (e.v)} - 218 \text{ S}$$

$$S = 0.511107$$

$$\text{PRESS} = 9.39081$$

$$r^2 = 87.7\%$$

3.5 Fourth QSAR model

MLR equation of this QSAR model P log 1/C is given by-

$$\text{Obsd log } 1/C = 65.1 - 8.47 \text{ E LUMO (e.v)} + 4.26 \text{ E HOMO (e.v)} - 539 \text{ S} + 0.0445 \text{ MR (cm}^3/\text{mol)}$$

$$S = 0.487954$$

$$\text{PRESS} = 8.71520$$

$$r^2 = 89.2\%$$

3.6 Fifth QSAR model

MLR equation of this QSAR model P log 1/C is given by-

$$\text{Obsd log } 1/C = 67.5 - 8.32 \text{ E LUMO (e.v)} + 4.54 \text{ E HOMO (e.v)} - 525 \text{ S}$$

$$- 0.0216 \text{ MR (cm}^3/\text{mol)} + 0.0212 \text{ MV (cm}^3/\text{mol)}$$

$$S = 0.472278$$

$$\text{PRESS} = 8.68476$$

$$r^2 = 90.2\%$$

3.7 Sixth QSAR model

MLR equation of this QSAR model P log 1/C is given by-

$$\text{Obsd log } 1/C = 15.0 - 5.43 \text{ E LUMO (e.v)} + 1.03 \text{ E HOMO (e.v)} - 127 \text{ S}$$

$$+ 0.183 \text{ MR (cm}^3/\text{mol)} + 0.0505 \text{ MV (cm}^3/\text{mol)} - 0.0379 \text{ Parachor (cm}^3/\text{mol)}$$

$$S = 0.437862$$

$$\text{PRESS} = 8.72525$$

$$r^2 = 91.9\%$$

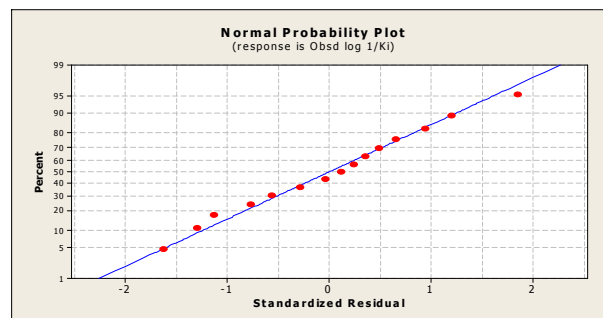


Figure 2: Correlation between observed and estimated log₁/k_i using model 6.

Table 2. Calculated values of quantum and physicochemical indices for the set of compounds used in the present study

Compd No.	Obsd log $1/K_i$	E LUMO (e.v)	E HOMO (e.v)	μ	η	S	χ	MR (cm ³ /mol)	MV (cm ³ /mol)	Parachor (cm ³ /mol)
1	8.600	-1.325	-9.021	-5.173	3.848	0.130	5.173	134.340	362.400	1032.000
2	8.460	-1.236	-8.964	-5.100	3.864	0.129	5.100	131.780	364.500	1023.500
3	8.400	-1.256	-9.011	-5.134	3.878	0.129	5.134	136.410	380.500	1063.600
4	7.750	-1.066	-8.960	-5.013	3.947	0.127	5.013	136.380	381.200	1061.500
5	9.100	-1.601	-8.999	-5.300	3.699	0.135	5.300	134.080	360.700	1042.600
6	8.510	-1.370	-9.053	-5.212	3.842	0.130	5.212	131.520	362.800	1034.000
7	8.850	-1.530	-9.105	-5.318	3.788	0.132	5.318	136.150	378.900	1074.100
8	8.680	-1.463	-9.106	-5.285	3.822	0.131	5.285	140.780	395.000	1114.200
9	8.230	-1.310	-9.070	-5.190	3.880	0.129	5.190	136.120	379.500	1072.000
10	8.820	-1.455	-8.858	-5.157	3.702	0.135	5.157	140.750	395.600	1112.100
11	8.510	-1.455	-9.214	-5.335	3.880	0.129	5.335	134.080	360.700	1042.600
12	8.680	-1.336	-9.065	-5.201	3.865	0.129	5.201	129.630	356.000	1002.300
13	8.220	-1.113	-8.811	-4.962	3.849	0.130	4.962	136.320	390.300	1073.800
14	7.960	-1.125	-9.020	-5.073	3.948	0.127	5.073	131.670	374.800	1031.700
15	8.550	-1.235	-8.899	-5.067	3.832	0.130	5.067	136.300	390.900	1071.800

4. CONCLUSIONS

Values of the descriptors of QSAR of HIV-1 Protease (RP) Inhibitors of Cycloalkylpyranone derivatives have been calculated using PM3 method and are given in table-2. With the help of these values of descriptors, six QSAR models have been developed using MLR analysis in different combinations of descriptors. The Chemical Potential (μ) and Absolute Hardness (η) descriptors have no predicting power and hence not included in the models. Best QSAR models is the model sixth listed below-

Sixth QSAR model

MLR equation of this QSAR model P log $1/C$ is given by-

$$\text{Obsd log } 1/C = 15.0 - 5.43 \text{ E LUMO (e.v)} + 1.03 \text{ E HOMO (e.v)} - 127 \text{ S}$$

$$+ 0.183 \text{ MR (cm}^3/\text{mol)} + 0.0505 \text{ MV (cm}^3/\text{mol)} - 0.0379 \text{ Parachor (cm}^3/\text{mol)}$$

$$S = 0.437862$$

$$\text{PRESS} = 8.72525$$

$$r^2 = 91.9\%$$

This is one of the best QSAR model in all the six models and has been developed using E LUMO, E HOMO, Global Softness (S), Molar refractivity (MR), Molar Volume (MV), Parachor (Pc). This MLR equation is given by Value of regression coefficient is 91.9% Prediction sum of squares coefficient (PRESS) is 8.72525 and Standard error of the regression (S) is 0.437862 which indicate the ability of predictive power of this QSAR model. QSAR model sixth can efficiently be used for the prediction of activity of any derivative of compound. The normal probability plot of responses is obsd log $1/C$ is shown in fig-2, which is clearly illustrates the high predictive power of the QSAR model six.

5. ACKNOWLEDGEMENT

We are thankful to the director of Central Drug Research Institute, Lucknow in order to provide library facilities.

REFERENCES

1. Hagen, S. E.; Prasad, J.V.N.V.; Boyer, F. E.; Domagala, J. M.; Ellsworth, E. L.; Gajda, C.; Hamilton, H. W.; Markoski, L. J.; Steinbaugh, B. A.; Tait, B. D.; Lunney, E. A.; Tummino, P. J.; Ferguson, D.; Hupe, D.; Nouhan, C.; Grachek, S. J.; Saunders, J. M. and Vander Rost S. *J. Med. Chem.* 1997, 40, 3707 (and the references therein).
2. (a) Popovic, M.; Sarngadharan, M. G.; Read, E.; Gallo, R. C. *Science* 1984, 224, 497. (b) Gallo, R. C.; Salahuddin, S. Z.; Popovic, M.; Shearer, G. M.; Kaplan, M.; Haynes, B. F.; Palker T. J.; Redfield, R.; Oleske, J.; Safai, B.; White, G.; Foster, P.; Markham, P. D. *Science* 1984, 224, 500.
3. Gallo, R. C.; Montagnier, L. *Sci. Am.* 1988, 259, 40.
- (4) Coffin, J.; Hasse, A.; Levy, J. A.; Montagnier, L.; Oroszlan, S.; Teich, N.; Temin, H.; Toyoshima, K.; Varmus, H.; Vogt, P.; Weiss, R. *Science* 1986, 232, 697.
5. Levy, J. A. *JAMA* 1989, 261, 2997 and references therein.
6. Clavel, F.; Guyader, M.; Guetard, D.; Salle, M.; Montagnier, L.; Alizon, M. *Nature (London)* 1986, 324, 691.
7. Guyader, M.; Emerman, M.; Sonigo, P.; Clavel, F.; Montagnier, L.; Alizon, M. *Nature (London)* 1987, 326, 662.
8. Rajni Garg, Satya P. Gupta, Hua Gao, Mekapati Suresh Babu, Asim Kumar Debnath, and Corwin Hansch, *Chem. Rev.* 1999, 99, 3525-3601.
9. R.P. Iczkowski, J.L. Margrave, *J. Am Soc.* 83, 354, 305.
- 10 R.S. Mulliken, *J. Chem. Phys.* 1934, 2, 782.
11. R.G. Parr.; Donnelly, R. A. Levy, Palke, W.E., *J. Chem. Soc.*, 1978, 68, 3801-3807.
12. Nelken, L. H., Index of Refraction, in *Handbook of Chemical Property Estimation*, W. J. Lyman, W. F. Reehl, and D. H. Rosenblatt, Editors, 1990. Washington, DC: American Chemical Society, p. 26.
13. Zhou, R.G.Parr, 1, *Am.Chem.Soc.*, 1990, (12), 5720
14. Grain, CR, Interfacial Tension with Water, in *Handbook of Chemical Property Estimation*, W. J. Lyman, W. R. Reehl, and D. H. Rosenblatt, Editors, 1990. Washington, DC: American Chemical Society.
15. Parr.; R.G., Yang, W. Density Function Theory of Atoms and molecules, Oxford University Press, New York, 1989.
16. W.Khon, A.D. Becke, R.G. Parr, *J. Chem. Phys.*, 1996, 100, 12974,.
17. F. A. Pasha, H. K. Srivastava, P. P. Singh, *Bioorganic & Medicinal Chemistry* 12, (2004) 171-177.
18. F. A. Pasha, H. K. Srivastava, P. P. Singh, *Combi. Sci.*, 2007, 26 (1), 69-84.
19. F. A. Pasha, Seung Joo Cho, Kaleem Ahmad, Yakub Beg, *Eur. J. of Med. Chem.* 43 2008, (17) 2361-2372.
20. P. P. Singh *et al*, *J. Chem. Pharm. Res.*, 2010, 2(5), 193-205.
21. Jignasa K. Savjani *et al*. *J. Chem. Pharm. Res.*, 2011, 3(1), 160-168.
22. P. P. Singh, Anil Kumar Gupta, Deepa Singh. *Chem Tech.*, 2011, 3(1), 165-177.
23. P.N, Tirpathi, Brajesh Singh, Kaleem Ahmad, Vivek Dixit. *SciTech*, 2011, 6(2), 23-27.
24. P. P. Singh, U. P. Singh and Dhruv Chandra Shukla, *J. Chem. Bio. Phy. Sci.* 2012, 6(2), 751-762.
25. Prabhat Kumar, Rajesh Kumar Pathak, P. P. Singh *J. Chem. Bio. Phy. Sci.*, 2012, 2(2), 612-621.